Calorimetry and Thermo Quiz ver1

Marks will be deducted for incorrect units or sig figs. A correct answer gets all points and work shown with an incorrect answer will likely get part marks

1. I tried to measure the heat energy produced by my stove by putting a 0.500kg aluminum pan (c=0.900J/g°C) on my stove. The frying pan heated up from 25°C to 229.0°C. How much energy did my stove produce? (/3)

2. In the previous class, Mr. Sheps burned a chip and a cheeto under a test tube and measured the change in temperature of the water. (/4)

1. What kind of reaction occurred (be specific): \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. Was it exothermic or endothermic: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. What type of a system was it? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. Describe this system in terms of energy and mass transfer:

3. A chemist took a 1.50kg block of iron (c=0.444J/g°C) and heated it up to 6.00x102K using 190.0 kilojoules of energy. What was the starting temperature (in Kelvin) of the block of iron? (/3)